Note Taking Guide: Episode 1102 **Name_____**

Acid/Base indicator: _____ that changes _____ in the presence of an _____ or a _____.

Solution	Litmus paper (R→B, B→R, or NC)	Phenolphthalein (color or NC)	Bromothymol Blue (color or NC)	Cabbage Juice (color or NC)
HNO_3				
NaOH				
KOH				
H_2SO_4				

Conclusion Questions:

1. Litmus turns _____ in an acid and _____ in a base.
2. Phenolphthalein (phth) turns _____ in an acid and _____ in a base.
3. Bromothymol blue turns _____ in an acid and _____ in a base.
4. Cabbage juice turns _____ in an acid and _____ in a base.

Strong Acids: dissociate _____ in _____ (ex: _____)
Weak Acids: dissociate _____ in _____
(ex: _____or _____)

Strong Bases: dissociate _____in _____ (ex: _____)
Weak Bases: dissociate _____in _____
(ex: _____or _____)

pH = _____

0_____7_____14

Determine the pH of a solution of HCl that has a molarity of 1×10^{-4} M.

Calculate the pH for a solution of HNO_3 with a molarity of _____.

Calculate the pH for a solution of H_2SO_4 with a molarity of _____.

$[H^+][OH^-]$ = _____

Calculate the pH of a solution of NaOH with a molarity of 3.0×10^{-2} M.

Find the pH for a solution of $Ca(OH)_2$ with a molarity of _____.

Calculate both the hydrogen ion concentration and the hydroxide ion concentration for an aqueous solution that has a pH of _____.

The Chemistry Quiz

CR1. CR2. 1. 2.

3. 4. 5.